**ASSSIGNMENT INFOGRAPHICS**

a) Use the standard electrode potentials for elements Q,R,S,T and U given below to

answer the question that follow the letters do not represent the actual symbol of the elements

Q2+(aq)+2e- Q1(s) -2.90

R2++(aq)+2e- R(s) - 2.38

S+(aq) +e- ½ S2(g)- 0.00

T2+(aq)+2e- T(s) + 0.34

½ U(g) + e- U-(aq) +2.87

a) Which element is likely to be hydrogen?. Give a reason for your answer(2mks)

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b) Which element is the weakest oxidizing agent? (1mk)

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c) i) In the space provided draw a well labeled diagram of the electrochemical

cell that would be obtained when half-cells of elements Q and T are combined (3mks)

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ii) Calculate the overall E value of the electrochemical cell constructed in

C (i) above (1mk)

d) Why is it advisable to a solution containing R2+ ions in a container made of T

(2mks)

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e) An iron spoon was being electroplated using silver. Draw a well labeled diagram to show how the iron spoon can be electroplated using silver. (2mks)